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for many gaseous solutes the relation between solubility c_g and partial pressure p_g is a proportional one $c_g = k p_g$ where k is a proportionality

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if p is a particular solution then $ap + b$ text and if x is a solution to the homogeneous equation $ax = 0$ text then $a(x + p) = ax + ap = 0 + b = b$ onumber so $x + p$ is another solution of $ax = b$ on the other hand if we start with any solution x to $ax = b$ then $x - p$ is a solution to $ax = 0$ since

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