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chemical calculations typical chemical calculations are studied in this chapter relationship between mass mole and number of particles percent composition by mass stoichiometry limiting reactant and yield calculation mass mole number relationship percent composition by mass stoichiometry limiting reactant percent yield

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chemical equations are symbolic representations of chemical and physical changes formulas for the substances undergoing the change reactants and substances generated by the change products are separated by an arrow and preceded by integer coefficients indicating their relative numbers

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chemists measure the properties of matter and express these measurements as quantities a quantity is an amount of something and consists of a number and a unit the number tells us how many or how much and the unit tells us what the scale of measurement is

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calculations involving conversions between moles of a material and the mass of that material are described

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chem1 chemical equations and calculations covers basic chemical arithmetic and stoichiometry for a course in general chemistry it is part of the general chemistry virtual textbook a free online reference textbook for general chemistry by stephen lower of simon fraser university

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in this article we ll look at how we can use the stoichiometric relationships contained in balanced chemical equations to determine amounts of substances consumed and produced in chemical reactions balanced equations and mole ratios

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4 1 writing and balancing chemical equations openstax

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extending this symbolism to represent both the identities and the relative quantities of substances undergoing a chemical or physical change involves writing and balancing a chemical equation

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chemical reactions introduction balancing chemical equations balancing more complex chemical equations visually understanding balancing chemical equations balancing another combustion reaction balancing chemical equation with substitution stoichiometry worked example calculating amounts of reactants and products

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a balanced chemical equation not only tells how many molecules of each kind are involved in a reaction it also indicates the amount of each substance that is involved the stoichiometric ratio measures one element or compound against another

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molarity or molar concentration is the number of moles of solute per liter of solution which can be calculated using the following equation molarity mol solute l of solution molar concentration can be used to convert between the mass or moles of solute and the volume of the solution introduction mixtures and solutions

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1 2 calculations the mole is the key concept for chemical calculations definition the mole is the amount of substance in grams that has the same number of particles as there are atoms in 12 grams of carbon 12 avogadro s number there are 6.022×10^{23} atoms in 12 grams of carbon 12 therefore explained in simpler terms one mole of any specified

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method write the formula of magnesium iodide by looking up their symbols on the periodic table magnesium is mg and iodine is i 1 to work out how many atoms of each element form a compound look at the charges that would form on each element

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a chemical formula is notation that shows the number and type of atoms in a molecule in other words it is a written representation of a three dimensional object there are a few different ways of writing a chemical formula in general a formula includes element symbols that identify the types of atoms in the molecule

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fischer projection lewis structure empirical formula structural formula molecular formula chemical formula any of several kinds of expressions of the composition or structure of chemical compounds the forms commonly encountered are empirical molecular structural and projection formulas

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a chemical equation shows the starting compound s the reactants on the left and the final compound s the products on the right separated by an arrow in a balanced chemical equation the numbers of atoms of each element and the total charge are the same on both sides of the equation

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the maximum mass of product formed in a reaction can be calculated using the balanced equation the mass of the limiting reactant and the ar relative atomic mass or mr relative formula

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word equations are useful to show which chemicals react together reactants and which chemicals are produced products symbol equations allow chemists to work out the masses that will

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