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worked example lewis diagram of formaldehyde CH_2O

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a covalent bond is formed between two atoms by sharing electrons the number of bonds an element forms in a covalent compound is determined by the number of electrons it needs to reach octet hydrogen is an exception to the octet rule h forms only one bond because it needs only two electrons

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covalent bond in chemistry the interatomic linkage that results from the sharing of an electron pair between two atoms the binding arises from the electrostatic attraction of their nuclei for the same electrons a bond forms when the bonded atoms have a lower total energy than that of widely separated atoms

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google classroom you might need periodic table complete the statements a single covalent bond involves shared electrons or shared pair s a double covalent bond involves shared electrons or shared pair s a triple covalent bond involves shared electrons or shared pair s

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multiple bonds first an atom may complete its octet by sharing more than one pair of electrons with a bonded neighbour two shared pairs of electrons represented by a double dash form a double bond double bonds are found in numerous compounds including carbon dioxide

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summary covalent bonds form when electrons are shared between atoms and are attracted by the nuclei of both atoms in pure covalent bonds the electrons are shared equally in polar covalent bonds the electrons are shared unequally as one atom exerts a stronger force of attraction on the electrons than the other

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