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## **atom definition structure history examples diagram *May 22 2024***

the single most important characteristic of an atom is its atomic number usually denoted by the letter z which is defined as the number of units of positive charge protons in the nucleus for example if an atom has a z of 6 it is carbon while a z of 92 corresponds to uranium

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all atoms of hydrogen have one and only one proton in the nucleus all atoms of iron have 26 protons in the nucleus this number of protons is so important to the identity of an atom that it is called the atomic number the number of protons in an atom is the atomic number of the element

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atomic structure is the focus of this session including a review of the famous backscattering experiment of rutherford

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the atomic structure refers to the structure of an atom comprising a nucleus centre in which the protons positively charged and neutrons neutral are present the negatively charged particles called electrons revolve around the centre of the nucleus

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definition according to dalton s theory atom is smallest particle which could not be divided any further atom is the entity that take part in a chemical reaction for example he and ne etc have atoms which exists independently while atoms of hydrogen nitrogen and oxygen do not exist independently

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summary the atom consists of discrete particles that govern its chemical and physical behavior each atom of an element contains the same number of protons which is the atomic number  $z$  neutral atoms have the same number of

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most of the mass of an atom is in the nucleus while the orbiting electrons account for an atom's size as a result an atom consists largely of empty space rutherford called his description the planetary model of the atom figure 2 2 2 2 2 2 shows how this model explains the experimental results

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