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stoichiometry is the calculation of the quantities of reactants or products in a chemical reaction using the relationships found in the balanced chemical equation the coefficients in a balanced chemical equation represent the reacting ratios of the substances in the reaction this page titled 12 stoichiometry is shared under a ck 12 license and was authored remixed and or curated by ck 12 foundation via source content that was edited to the style and standards of the libretxts platform a detailed edit history is available upon request stoichiometry problems when doing stoichiometry problems you follow 3 steps step 1 convert to moles step 2 mole ratio step 3 convert to unit asked for sometimes you can skip one or two steps section 12 2 chemical calculations objectives calculate stoichiometric quantities from balanced chemical equations using units of moles mass representative particles and volumes of gases at stp mole to mole conversions $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$ each time we use 2 moles of Al_2O_3 we will also make 3 moles of O_2 mole to mole conversions how many 12 1 stoichiometry review we have used balanced equations to set up ratios in terms of moles of materials that we can use as conversion factors to answer stoichiometric questions such as how many moles of substance a react with so many moles of reactant b 6 12 6 what volume of carbon dioxide is produced when 3 7 l of oxygen are consumed the compound tnt trinitrotoluene decomposes explosively into carbon carbon monoxide hydrogen and nitrogen what volumes of hydrogen and nitrogen are produced if 5 8 l of co is produced the balanced equation is $2\text{C}_6\text{H}_6 + 12\text{O}_2 \rightarrow 5\text{H}_2\text{O} + 3\text{N}_2$ 17 how many stoichiometry this section explains how to calculate the amount of reactants required or product formed in a nonchemical process it teaches you how to interpret chemical equations in terms of interacting moles representative particles masses and gas volume at stp 1 the quantitative relationship among reactants and products is called stoichiometry the term stoichiometry is derived from two greek words stoicheion meaning element and metron meaning measure on this subject you often are required to calculate quantities of reactants or products chapter 12 stoichiometry 12 1 the arithmetic of equations 12 1 lesson check page 389 7 answer a balanced chemical equation provides the same kind of quantitative information that a recipe does you will calculate the number of moles and the mass of a reactant or product when given the number of moles or the mass of another reactant or product you will identify the limiting reactant in a chemical compounds with integral atomic ratios like nitrous oxide are described as stoichiometric compounds and they permit many simple calculations the common oxide of aluminum provides a second example but this time begin with the mass percent and deduce the atomic ratio 1 the document provides a study guide for stoichiometry concepts including mole ratios limiting reagents theoretical yield actual yield and percent yield 2 it includes example stoichiometry problems and their step by step worked solutions through online tutorials this chapter will describe how to symbolize chemical reactions using chemical equations how to classify some common chemical reactions by identifying patterns of reactivity and how to determine the quantitative relations between the amounts of substances involved in chemical reactions that is the reaction stoichiometry chapter 12 review stoichiometry important vocabulary stoichiometry the calculations of quantities in chemical reactions in a subject of chemistry mole ratio a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles for the basics of stoichiometry review this chapter excluding the stoichiometry applications as we will be looking at those in more detail later in this chapter for 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