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8 5 redox reactions the oxidation number of an element represents the total number of electrons which have been removed a positive oxidation state or added a negative oxidation state to get the element into its present state

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oxidation reduction reactions commonly known as redox reactions are reactions that involve the transfer of electrons from one species to another the species that loses electrons is said to be oxidized while the species that gains electrons is said to be reduced

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learn what a redox reaction is and how to identify and balance redox reactions explore oxidation reduction and oxidation numbers

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what are redox reactions redox is a shorthand for reduction oxidation meaning that a redox reaction is one in which both a reduction reaction and an oxidation reaction takes place at once it is also shorthand for oxidation reduction reaction

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a redox reaction is a chemical reaction in which the atoms change their oxidation numbers some atoms lose electrons and are oxidized a process known as oxidation on the other hand some atoms gain electrons and are reduced a process known as reduction therefore both reduction and oxidation take place simultaneously hence the term redox

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the terms reduction and oxidation are usually abbreviated to redox such a reaction corresponds to the transfer of electrons from one species to another a simple redox reaction occurs when copper metal is immersed in a solution of silver nitrate

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for balancing redox reactions it is necessary to first balance the main atoms through adjusting stoichiometric coefficients then the charges through electron transfer and as per conditions acidic neutral basic

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why study redox reduction oxidation reactions 1 redox reactions fuel and constrain almost all life processes 2 redox reactions are a major determinant of chemical species present in natural environments redox reactions are characterized by the transfer of electrons between chemical species

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redox reactions oxidation and reduction take place together at the same time in the same reaction these are called redox reactions oxidation is a reaction in which oxygen is added to an element or a compound reduction is a reaction in which oxygen is removed from an element or compound

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redox reactions are oxidation reduction reactions where reactants undergo changes in oxidation states understand balancing redox reactions with examples types and applications

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redox reactions redox reactions are reactions in which oxidation and reduction take place together while one species is oxidising another is reducing in the same reaction eg $\text{Cu} + 2\text{Mg} \rightarrow \text{Cu} + 2\text{Mg}^{2+}$ Cu has been reduced and Mg has been oxidised

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a redox is a type of chemical reaction that involves a transfer of electrons between two species the oxidation state of an element corresponds to the number of electrons that an atom loses gains in a chemical reaction the reducing agent is the reactant that is being oxidized the oxidizing agent is reactant that is being reduced

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oxidation reduction reactions also called redox reactions involve the transfer of electrons from one species to another that electron transfer causes a change in oxidation state for both reactive partners

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redox reactions and electrode processes when zinc rod is dipped in copper sulphate solution redox reaction begins hence zinc is oxidised to Zn^{2+} ions and Cu^{2+} ions are reduced to metal redox reaction reactions in which oxidation and reduction

occur simultaneously are called redox reactions oxidation involves loss of one or more

lecture 25 oxidation reduction and electrochemical cells *Dec 10 2022*

introduction to oxidation reduction redox reactions balancing redox reactions electrochemical cells lecture video redox reactions are a major class of chemical reactions in which there is an exchange of electrons from one species to another in this lecture the basic principles of redox reactions are introduced

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